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An Arrhenius

acid is any

substance that

produces

hydrogen ions

[or hydronium

ions (H_3O^+) a

hydrogen ion

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attached to a water molecule] when dissolved in water. An Arrhenius base is any substance that produces hydroxide ions when dissolved in water. While the Arrhenius definitions of acids and bases is useful, it is

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we encounter acids. such as vinegar and lemon juice and bases such as ammonia and lye in drain cleaners. Many important processes in chemistry and biology involve reactions of acids with

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In some

textbooks, when explaining the Brønsted – Lowry definition, acids and bases are described as proton donors and proton acceptors. Based on your response to Exercise 2,

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explain why

these are
correct terms.

ion desca\oed

(1B 4. A

definition of
the prefix amphi
is; "both or of
both kinds."

Mrs.

Zuberbuehler -

Mrs.

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STOP Conjugate

Acid $\text{CO}_3^{2-}(\text{aq}) +$

$\text{H}^+(\text{aq})$ Model

3 - Conjugate

Acid-Base Pairs

Base + 1-120 (1)

$\text{HCO}_3^-(\text{aq})$ 13.

All acid-base

reactions have

two conjugate

acid-base pairs.

One conjugate

acid-base pair

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in Model 3 is $\text{H}^+/\text{H}_2\text{O}$. List the other acid-base pair in the reaction. 14.

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Phosphoric acid
(H_3PO_4) is a
triprotic acid,
but phosphorous
acid (H_3PO_3)
is a diprotic
acid, and
hypophosphorous
acid (H_3PO_2)
is a monoprotic
acid. a. Draw
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diagrams for
each of these
three acids. All
of the
structures
include a double
bonded oxygen.

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