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Chemistry (Specification 2420) CHEM5 Unit 5: Energetics,
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Mark Scheme – General Certificate of Education (A-level) Chemistry – Unit 5: Energetics, Redox and Inorganic Chemistry – January 2011 6 2(g) $H = 1.9 \text{ (kJ mol}^{-1}\text{)}$
 $S = 2.4 - 5.7 = -3.3 \text{ (J K}^{-1} \text{ mol}^{-1}\text{)}$ G is always positive 1 1 1 for M1 and M2 allow no units, penalise wrong units

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...

5(c) Any ethanol present in the breath of a drinker can be detected by using a breathalyser. The ethanol is converted into ethanoic acid. The breathalyser has negative and positive electrodes. A current is measured and displayed in terms of alcohol content. The overall redox equation is as follows $\text{CH}_3\text{CH}_2\text{OH(l)} + \text{O}_2\text{(g)} \rightarrow \text{CH}_3\text{COOH(l)} + \text{H}_2\text{O(l)}$

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- (a) $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ octahedral 1 1 Only allow if species has 6 ligands but allow if M1 not given because charge missing 7
(b) CoCO_3 Purple solid (allow pink) 1 1 Mark independently

~~A-level Chemistry Mark Scheme January 2010~~

January 2013 QP - Unit 1 Edexcel Chemistry A-level January 2014 (IAL) MS - Unit 1 Edexcel Chemistry A-level January 2014 (IAL) QP - Unit 1 Edexcel Chemistry A-level

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