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(i) Consider the rate law  $v = k[H_2][I_2]$ . If we substitute units into the equation, we obtain  $(\text{mol dm}^{-3} \text{ s}^{-1}) = [k] (\text{mol dm}^{-3}) (\text{mol dm}^{-3})$  where the notation  $[k]$  means 'the units of  $k$ '. We can rearrange this expression to find the units of the rate constant,  $k$ .  $[k] = (\text{mol dm}^{-3} \text{ s}^{-1}) (\text{mol dm}^{-3}) (\text{mol dm}^{-3}) = \text{mol}^{-1} \text{ dm}^3 \text{ s}^{-1}$

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Chemical Kinetics, 3Rd Edition by Keith J. Laidler ISBN 13: 9788131709726 ISBN 10: 8131709728 Paperback; New Delhi: Pearson India, 2003; ISBN-13: 978-8131709726

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Chemical Kinetics by Laidler, Keith J. Tata McGraw-Hill Publishing, 1976. Soft cover. Very Good. Published 1976. p/b 566 pages, condition is very good, minor edge wear. ...

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