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Electrode, electrode
potential \u0026
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Chapter 20
(Electrochemistry) -
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Part 1

Chapter 20 –
Electrochemistry: Part
1 of 13

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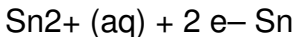
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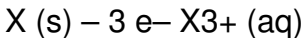
Electrochemistry -

Questions With
Answers Answer: (a)

tin electrode is the cathode; cathode is the site of reduction (gain in electrons) and will convert metal ions into a metal. (b) (see diagram) (c) red:



(s) $E^{\circ} = -0.14 \text{ V}$ oxid:



$E^{\circ} = +0.74 \text{ V}$ $E^{\circ} \text{ cell} =$

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+0.60 V red: $X^{3+} (aq)$
 $+ 3 e^{-} \rightarrow X$

~~AP REVIEW~~

~~QUESTIONS~~

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~~Answers~~

Class 9 Chemistry -

Chapter 7 -

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Easy notes that
contain all exercise
questions of the

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chapter.

Questions With

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~~Review Questions~~

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Electrochemistry

Review Practice

Questions . Multiple

Choice . 1. If a neutral

atom becomes

positively charged

then it has a. been

reduced. b. been

oxidized. c. lost a

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proton. d. gained a
neutron. 2. Which
change would be
classified as an
oxidation? a) $\text{Cr}^{2+} \rightarrow \text{Cr}^{3+}$?
b) $2\text{H}^+ \rightarrow \text{H}_2$? c)
 $\text{Br}_2 \rightarrow 2\text{Br}^-$? d) $\text{Mg}^{2+} \rightarrow \text{Mg}$. 3. When the half-
reaction $\text{Cr}^{2+} + 7\text{H}_2\text{O} \rightarrow \text{Cr}^{3+} + 14\text{H}^+$

~~Chapter 14~~
~~Electrochemistry~~
~~Review Practice~~

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Questions

Electrochemistry
Review Questions. In
the reaction: $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$. the
element that
undergoes reduction
is: none - this is not a
redox reaction.
oxygen. sodium.
sulfur. hydrogen.

Electrochemistry

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Chemistry 30 - Redox

and Electrochemistry

Review Test pag e 4

a) Cd (s) b) Co (s) c)

Pb (s) d) Cu (s) Use

the following

information to answer

question 18. A

student observed the

reactions between

four different metals

and the solutions of

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their ions, and then

recorded these
“spontaneous”

reactions: $W(s) + X$

$+ (aq) \rightarrow W + (aq) + X$

$(s) \parallel X(s) + Y + (aq) \rightarrow$

$X +$

~~Chemistry 30 Review~~

~~Test 3 – Ms. Mogek's~~

~~Classroom~~

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Review Questions.

Choose the correct

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Review for each
question. Show all
questions $\Leftarrow \Rightarrow$ The
Answers
oxidation number on
nitrogen in a nitrate
ion is: ? -1 ? -3 ? -5 ?
+3 ? +5; The sum of
the oxidation numbers
on all of the atoms in
the formula of an
acetate ion is equal
to: ? -2 ? -1 ...

Electrochemistry

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Examination

Questions — Exam

Answers Free

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37162; Contributors

and Attributions;

Chemical reactions

involving transfer of

electrons are called

oxidation and

reduction reactions or

redox reactions. When

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properly set up, these reactions generate power in a Battery, or galvanic cell.

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5. Tying

Electrochemistry to Thermodynamics. In electrochemistry, the quantity in which we are most interested is

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E, the potential energy of the system. It is the value you see on a new $E = 1.5\text{V}$ or $E = 6\text{ V}$ battery. We can relate this idea of work done in electrochemistry to the thermodynamic concept of work, free energy, through the equation:

~~Chapter 21: ELECTR~~

Page 23/36

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~~OCHEMISTRY~~

~~TYING IT ALL~~

~~TOGETHER~~

AP Chemistry Review
Questions -

Electrochemistry. If a constant current of 8.00 amperes is passed through a cell containing Zn^{2+} for 2.00 hours, how many grams of zinc will plate out onto the cathode? 0.985 g.

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1.43 x 10³ g. 39.0 g.

126 g. 19.5 g.

~~AP Chemistry Review
Questions~~

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Electrochemistry

Practice Test with

Answers. Question 1.

On passing 0.5

faraday of electricity

through NaCl, the

amount of Cl

deposited on cathode

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is. Related: PES
Scholastic Aptitude
Test Chemistry. A.
17.75 gm. B.

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Multiple Choice
Questions Answers—
JEE ...~~

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Diploma 3; Diploma 4;
Diploma 5; Diploma 6;

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Diploma 7; Diploma 8;
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Review; Formula ...

Questions With
~~Diploma Practice With
Answers~~ WG

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Questions 6-9 are related to the distinction between adsorbed and diffusive redox species. The answers to questions 10-13 explain the interpretation of slow

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and fast scan

voltammetry with
redox proteins.

Questions 14-19 deal
with catalytic
electrochemistry,
when the protein
studied is actually an
enzyme. Questions
20, 21 and 22 are
general.

Protein

Electrochemistry:

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Questions and
Answers

the answer. (2)

SOLUTIONS TO ELE
CTROCHEMISTRY

Question 1 1.1

Pressure: 101,3 kPa

(1,013 x 10⁵ Pa)

Temperature: 25 °C

(298 K) 1.2 Salt

Bridge 1.3 Anode ,

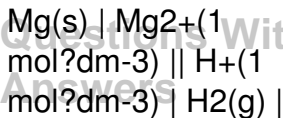
Mg is a stronger

reducing agent than

H₂ and therefore (Mg)

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will be oxidised. 1.4



Pt(s) 1.5 $E_{\text{cell}} = E_{\text{cathode}} - E_{\text{anode}}$

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~~Mindset Learn~~

The electrochemical series shows equations representing reactions of which type? 2.

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Which of the following metals is highest in the electrochemical series? 3. If a cell is set up with copper and one other metal, which other metal would allow a flow of electrons AWAY FROM the copper? 4.

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oxidation number on
nitrogen in a nitrate
ion is 1 3 5 3 5 the
sum of the oxidation
numbers on all of the
atoms in the formula
of an acetate ion is
equal to 2 1 question
6 total score 9

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ed4a5a08ab6ad2c26

bae48e336a6 Questions With

Answers