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Molarity Dilution Problems Solution

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molarity =  $\frac{\text{mol solute}}{\text{L solution}}$ . 1 L = 1000 mL. The molarity of a solution is a ratio of the moles of solute per liters of solution. The units for molarity are written as mol/L or M. This measurement is used to perform stoichiometric calculations.

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ml. A solution can be made less concentrated

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in a process called dilution. This is accomplished by adding more solvent.

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EQUATIONS molarity =  $\frac{\text{L. solution mol solute}}{\text{L}}$  1  
L = 1000 mL The molarity of a solution is a  
ratio of the moles of.

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Solution Stoichiometry Worksheet Solve the  
following solutions Stoichiometry problems:

1. How many grams of silver chromate will  
precipitate when 150. mL of 0.500 M silver  
nitrate are added to 100. mL of 0.400 M  
potassium chromate?  $2 \text{ AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{ KNO}_3(\text{aq})$  0.150 L  
 $\text{AgNO}_3$  0.500 moles  $\text{AgNO}_3$  1 moles  $\text{Ag}_2\text{CrO}_4$

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As we learned previously, double replacement reactions involve the reaction between ionic

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compounds in solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because these reactions occur in aqueous solution, we can use the concept of molarity to directly calculate the number of moles of reactants or products that will ...

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calcium phosphate can be produced from the  
reaction of 2.50 L of 0.250 M Calcium  
chloride with and excess of phosphoric acid?

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The stoichiometry of chemical reactions may serve as the basis for quantitative chemical analysis methods. Titrations involve measuring the volume of a titrant solution required to completely react with a sample solution. This volume is then used to calculate the concentration of analyte in the sample using the stoichiometry of the titration ...

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