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Mass /u0026amp; Volume

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Liters Using Molarity -

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Molarity as a

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Mole, Grams to

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Worksheet problems: 1. How

many grams of silver chromate will

precipitate when 150.

mL of 0.500 M silver

nitrate are added to

100. mL of 0.400 M

potassium chromate?

$2 \text{AgNO}_3(\text{aq}) + \text{K}_2$

$\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{KNO}_3(\text{aq})$

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Solution

0.150 L AgNO_3 0.500

moles AgNO_3 1

moles Ag_2CrO_4 4

331.74 g Ag_2CrO_4 =

12.4 g Ag_2CrO_4 1 L

2 moles ...

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Chemistry 11

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Directions: Answer in
the

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Worksheet :

Stoichiometry (using
solutions) 1. Given

the following

reaction: (hint:

balance the equation

first) $\text{H}_2\text{SO}_4 + \text{NaOH}$

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g Na₂SO₄ + H₂O. If
43.2 mL of 0.236 M
NaOH reacts with
36.7 mL of H₂SO₄,
what is the
concentration of the
H₂SO₄ solution?
answer. 2. Given the
following equation:
NaOH + HCl → H₂O +
NaCl. If 36.7 mL of
HCl solution is
needed to react with
43.2 mL of a 0 ...

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~~Stoichiometry (using
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Stoichiometry Using
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Solutions

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Solutions Worksheet :
Stoichiometry (using
solutions) 1. Given
the following
reaction: (hint:

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balance the equation
first) $\text{H}_2\text{SO}_4 + \text{NaOH}$
g Na_2SO_4 ...

Calculate the
molarity of the H_2SO_4
solution if it takes
40.0 mL of H_2SO_4 to
neutralize 0.364 g of
 Na_2CO_3 . answer ...
Molarity (M) Solution

...

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~~Solutions~~ ~~Stoichiometry~~

Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.

$1.54 \text{ mol K}_2\text{SO}_4 = 12.3 \text{ M K}_2\text{SO}_4 \text{ } 0.125 \text{ L soln}$
A chalice contains 36.45 grams...

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Because these
reactions occur in

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aqueous solution, we can use the concept of molarity to directly calculate the number of moles of reactants or products that will be formed, and hence their amounts (i.e. volume of solutions or mass of precipitates). As an example, lead(II) nitrate and sodium chloride react to form

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sodium nitrate and
the insoluble
compound, lead(II)
chloride.



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molarity of the

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following solutions
given that: 1) 1.0
moles of potassium
fluoride is dissolved
to make 0.10 L of
solution.

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molarity = $\frac{\text{mol solute}}{\text{L solution}}$
1 L = 1000 mL
The molarity of a
solution is a ratio of

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the moles of solute per liters of solution. The units for molarity are written as mol/L or M. This measurement is used to perform stoichiometric calculations. The strategy used for solving these solution stoichiometry problems is to set up

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the problem so that
the units cancel.
When the volume of
a solution ...

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molarity, and then a
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Answers

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Answers Answers: 1a.

11.0 L of 0.5 M

$\text{Ca}(\text{OH})_2$ (aq) 3a. 107

g NH_4Cl 1b. 5.5 mol

H_2SO_4 3b. 296 g

$\text{Ca}(\text{OH})_2$ 1c. 6.71 L of

0.82 M H_2SO_4 (aq)

3c. 111 g CaCl_2 2a.

204 g CaCO_3 4a. 7.95

L of 0.1 M HCl (aq) 2b.

1.36 L of 3.0 M

HCl (aq) 4b. 16.4 g Zn

2c. 2.04 mol ...

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~~Finding Molarity,~~

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key - Bing As we

learned previously,

double replacement

reactions involve the

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reaction between ionic compounds in solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because ...

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comprehensive
collection of solved
exercises and

problems typical to

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